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Chapter 9 Chemical Names and Formulas 83 SECTION 9.3 NAMING AND WRITING FORMULAS FOR MOLECULAR COMPOUNDS (pages 268 – 270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268 – 269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

Name Date Class CHEMICAL NAMES AND FORMULAS 9

SECTION 9.2 NAMING AND WRITING FORMULAS FOR IONIC COMPOUNDS I.

Write the formulas for these binary ionic compounds. c. potassium iodide 2. write the formulas for the compounds formed c. b. f. 3. Name the following binary ionic compounds. sodium sulfide g. CuCl: h. SnCl_2 i. a. MnO_2 c. d. SrBr_2 221 Chapter 9 Chemical Names and Formulas

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1. Chapter 9 “ Chemical Names and Formulas ” H_2O 1. 2. Section 9.1 Naming Ions OBJECTIVES: • Identify the charges on monatomic ions by using the periodic table, and name the ions. • Define a polyatomic ion and write the names and formulas of the most common polyatomic ions. • Identify the two common endings for the names of most polyatomic ions. Atoms and Ions • Atoms are electrically neutral.

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CHEMISTRY – Chapter 9 – Scotch Plains-Fanwood High School Page 1 Chapters 9 -
Chemical Names/Formulas IONS = Charged particles conduct e-in solution = (aq)
CATIONS = + ANIONS = - Ionic Compounds – conduct electricity when (aq) or (l)
Many [ionic] formula units all attached. ex. NaCl Molecular Compounds – do NOT
conduct electricity Individual distinct molecules. ex. H₂O

Chapter 9 Packet - Chapters 9 Chemical Names/Formulas ...

Chapter 9 Practice Test - Naming and Writing Chemical Formulas Answer Section
MATCHING 1. ANS: A PTS: 1 DIF: L1 REF: p. 253 OBJ: 9.1.1 Identify the charges
of monatomic ions by using the periodic table, and name the ions. STA: Ch.5.a 2.ANS:
H PTS: 1 DIF: L1 REF: p. 254

Chapter 9 Practice Test - Naming and Writing Chemical Formulas

The name of a base is the name of the cation followed by the name of the anion
(hydroxide). The formula of a base is written by showing the number of hydroxide
ions needed to balance the positive charge on the cation. After reading Lesson 9.4,
answer the following questions. Names and Formulas of Acids hydrogen1.

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Bishop's text shows students how to break the material of preparatory chemistry down and master it. The system of objectives tells the students exactly what they must learn in each chapter and where to find it.

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reaction, dietary fiber, fat mimetics, and polyols, in addition to discussions of physical properties, imparted functionalities, and actual applications. It is an invaluable resource on the chemistry of food carbohydrates for advanced undergraduate and graduate students, and a concise, user-friendly, applied reference book for food science professionals. Identifies structures and chemistry of all food carbohydrates – monosaccharides, oligosaccharides and polysaccharides Covers the behavior and functionality of carbohydrates within foods Contains extensive coverage of the structures and properties of individual polysaccharides, including cellulose, inulin, gellans and pectins, amongst others

For the first time in over twenty-five years, this unique and popular textbook on food chemistry mechanism and theory has received a full update. Emphasizing the underlying chemical reactions and interactions that occur in foods during processing and storage, this book unifies the themes of "what", "how" and "why" in the language of equations, reactions and mechanisms. This book is the only work which provides in-depth focus on aspects of reaction mechanisms and theories in the chemistry of food and food systems. With more than 500 chemical equations and figures, this book provides unusual clarity and relevance, and fills a significant gap in food chemistry literature. It is a definitive source to consult regarding the important mechanisms that make food components and reactions tick. Mechanism and Theory in Food Chemistry has been a popular resource for students and researchers alike since its publication in 1989. This important new edition contains updates on the original text

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encompassing a quarter century of advances in food chemistry. Many parts of the original chapters are revised to make for smoother navigation through the subjects, to better explain the underlying chemistry concepts and to fulfill the need of adding topics of emerging importance. New sections on fatty acids, lipid oxidation, meat, milk, soybean and wheat proteins, starch and many more have been incorporated throughout the revision. This updated edition provides an excellent source of all the important chemical mechanisms and theories involved with food science.

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